



Standard Test Method for Saponification Value of Drying Oils, Fatty Acids, and Polymerized Fatty Acids¹

This standard is issued under the fixed designation D 1962; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

This standard has been approved for use by agencies of the Department of Defense.

1. Scope

1.1 This test method covers the determination of the saponification value of drying oils, bodied oils, fatty acids, and polymerized fatty acids.

1.2 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* For specific hazard statements, see Section 7.

2. Referenced Documents

2.1 ASTM Standards:

- D 305 Test Method for Solvent-Extractable Material in Black Pigments²
- D 1193 Specification for Reagent Water³

3. Terminology

3.1 Definitions:

3.1.1 *saponification value*—a measure of the alkali reactive groups in oils and fatty acids and is expressed as the number of milligrams of potassium hydroxide that react with 1 g of sample.

4. Significance and Use

4.1 The saponification value of oils and fatty acids is a measure of the content of ester linkages. For an oil, provided it is not significantly oxidized, the number of ester linkages per molecule (for example, three in a triglyceride), can be used to calculate the molecular weight of the oil.

4.2 A saponification value higher than normal indicates that the oil has been oxidized (blown) or chemically modified, for example, with other acids such as maleic, fumaric, or phthalic.

4.3 Saponification value alone is not a measure of the quality of the oil.

4.4 Chemically modified oils may require saponification

times longer than 1 h for complete reaction.

5. Apparatus

5.1 *Erlenmeyer Flasks*, wide-mouth, alkali-resistant, 250 or 300-mL capacity.

5.2 *Condenser Loop*.

NOTE 1—Suitable condenser loops are shown in Figs. 1 and 2 of Test Method D 305.

5.3 *Steam Bath*.

6. Reagents

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the Specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁴ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Type I of Specification D 1193.

6.3 *Phenolphthalein Indicator Solution*—Dissolve 1 g of phenolphthalein in 100 mL of ethanol (95 %), methanol or isopropanol.

NOTE 2—A “masked phenolphthalein indicator” may be used with off-color materials. Prepare by dissolving 1.6 g of phenolphthalein and 2.7 g of methylene blue in 500 mL of alcohol. Adjust the pH with sodium hydroxide (NaOH) or KOH solution so that the greenish blue color is faintly tinged with purple. The color change is from green to purple when going from acid to alkali.

6.4 *Potassium Hydroxide, Alcoholic Solution*—Place 5 to 10 g of potassium hydroxide (KOH) (**Caution**—see 7.1) in a 2-L flask and add 1 to 1.5 L of ethyl alcohol (95 %) or denatured alcohol conforming to Formula No. 30 or 3A of the U. S. Bureau of Alcohol, Tobacco and Firearms. Boil on a

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² *Annual Book of ASTM Standards*, Vol 06.03.

³ *Annual Book of ASTM Standards*, Vol 11.01.

⁴ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.